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**Include file name:** Chemistry\_Worksheet\_0016

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1. From the periodic table give the number of electrons, protons, and neutron for the most common isotope of Ca. (Give answer like this 4,4,4)
2. From the periodic table give the number of electrons, protons, and neutron for the most common isotope of Ir. (Give answer like this 4,4,4)
3. From the periodic table give the number of electrons, protons, and neutron for the most common isotope of Au. (Give answer like this 4,4,4)
4. From the periodic table give the number of electrons, protons, and neutron for the most common isotope of Pr. (Give answer like this 4,4,4)
5. An element has 3 isotopes one is 78.99% @ 23.99amu, the second is 10.00% @ 24.99amu and the last is 25.98amu. What is the atomic weight of this element. Report in significant figures and report number and unit with no space between them
6. Cl is in what period?
7. Mg is in what series
8. Cl is in what family
9. Ge is classified as a(n) (not family name)
10. Xe is in what family
11. Another name for noble gases is
12. Br is called a(n)(not family name)
13. Ca is classified as a(not family name)
14. The max electrons in 4th energy level is
15. Name the four subshells in an energy level (report with a coma and space to separate ex. 3, 3, 3, 3)
16. Write the electron configuration for P (leave no spaces between letters and numbers)
17. How many dots in a Lewis Dot does Fr have?
18. How many dots does Ar have in the Lewis Dot?
1. How many dots does Br have in the Lewis Dot?
2. How many electrons, protons & neutrons does Fr have? List e, p, n.
3. If an element has 29 electrons and 35 neutrons what is mass/
4. I is in what family?
5. He is classified as a? Nonmetal
6. Cs is in what series?

7. Lr is in what period?
8.  $1s^2 2s^2 2p^6 3s^2 4s^1$  give symbol.
9.  $1s^2 2s^2 2p^6 3s^2 2p^5$  is what symbol?
10. If an element has 23 electrons and 56 neutrons what is mass?

1. The concept of quantized energy level in an atom was the first proposed by:  
Democritus

Marchette

Bohr

Lavoiser

2. Dalton's atomic theory does not include which the following statements:

All atoms of the same element are identical to each other.

All atoms are colorless.

Atoms combine to form molecules. The molecule is a tightly bound group of atoms that acts as a unit.

All atoms of the same element are identical to each other. b. All atoms are colorless. c. Atoms combine to form molecules. The molecule is a tightly bound group of atoms that acts as a unit. d. All matter is made up of tiny indivisible particle called atoms

3. The mass number of an atom that has 20 neutrons and a charge of +18 in the nucleus is:

38

can not determined

56

20

4. The atomic number of the atom with 20 neutrons, and a charge of +18 in the nucleus is:

can not determine

20

38

18

5. The ground state electron configuration for an atom which has 6 protons and 6 neutrons in the nucleus:

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

6. A possible excited state electronic configuration for the atom in the above question is:

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \downarrow$

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

$\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

7. Which of the following elements is a metalloid?

Ag

Ge

He

H

8. Which of the following group of elements are nonmetals?

O, S, Se

. F, Cl, Al

Cu, Ag, Au

Na, K, Rb

9. The four most abundant elements in the earth's crust, in increasing order is:

Cl < O < N < P

Ca < F < O < Si

Fe < Al < Si < O

N > C > H > O

10. The mass of a single atom of lead is:

3.5G

35000000

.0000000000000000000035 g

3.5ng

11. The four most abundant elements in the human body are:

C, N, S, H

C, N, O, H

H, Li, C, N

Fe, C, N, Na

12. The four most abundant elements in the human body are considered:

Metalloids

transition metals

metals

representative metals

13. The two most abundant elements in the crust are:

H, He

O, Si

H, C

Ag, Au

14. An example of periodicity can be seen in:

Li, Be, C, N

H, He, B, C

H, Li, Na, K

C, N, Al, Si

15. Which is not an elementary particle of an atom:

protons

electron

pistron

neutron

16. Which Alkali metal takes the least energy to ionize an electron off:

Cs

Al

Na

K

17. The atomic nucleus has a \_\_\_\_ charge:

depends on state

positive

neutral

negative

18. Isotopes have a different number of:

electrons

neutrons

pistrons

protons

19. A positive ion is produced from an atom by:

Gaining electron

loosing electron

Gaining neutron

loosing proton

20. Ca is a \_\_\_\_ element:

metalloid

Alkaline metal

non metal

Alkali metal

21. Ionization energy in elements 1A \_\_\_\_ going down

can not determine

increases

decreases

is constant

22. Elements in the same \_\_\_\_ have the same valence electron

classification

Series

period

family

23. Group VIIA are called:

metals

alkali metals

halogens

noble gases

24. Which does not apply to O:

non metal

halogen

p5

. -2 ion

25. Which does not apply to Kr:

-1 ion

non metal

p6

.inert gas

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