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001 (part 1 of 1) 10 points

What is the formula for aluminum chloride?

1. AlCl_3
2. AlCl_2
3. AlCl
4. Al_3Cl
5. Al_2Cl

002 (part 1 of 1) 10 points

Cobalt (III) oxide would be the label for

1. Co_2O_3 .
2. Co_3O .
3. CoO .
4. CoO_3 .
5. Co_3O_2 .

003 (part 1 of 1) 10 points

Give the formula for sodium nitrate.

1. $\text{Na}(\text{NO}_3)_2$
2. $\text{Na}(\text{NO})_3$
3. Na_3NO_3
4. NaNO_3

004 (part 1 of 1) 10 points

How many hydrogen atoms are in a formula unit of ammonium nitrate? NH_4NO_3

1. 6 atoms
2. 8 atoms
3. 4 atoms
4. 2 atoms
5. 10 atoms

005 (part 1 of 1) 10 points

Name the compound K_2CO_3 .

1. potassium(II) carbonate
2. potassium carbonate
3. potassium carboxide
4. potassium carbide

006 (part 1 of 1) 10 points

Of the following combinations of compounds and names, the one which is incorrect is

1. Cu_2O : copper(I) oxide
2. LiCl : lithium chloride
3. Cs_2S : cesium sulfide
4. SO_3 : sulfur trioxide
5. N_2O_5 : dinitrogen pentoxide

007 (part 1 of 1) 10 points

What is the correct formula for zinc phosphate?

1. $\text{Zn}_2(\text{PO}_4)_3$
2. ZnPO_4
3. ZnP
4. $\text{Zn}_3(\text{PO}_4)_2$
5. $\text{Zn}(\text{PO}_3)_2$

008 (part 1 of 1) 10 points

How many oxygen atoms are in the formula for iron(III) oxide? Fe_2O_3

1. 1
2. 4
3. 3
4. 2

009 (part 1 of 1) 10 points

The binary compound PCl_3 is called

1. monophosphorus trichloride.
2. triphosphorus chloride.
3. phosphorus trichloride.
4. None of these.
5. phosphorus chloride.

010 (part 1 of 1) 10 points

What is the name for the compound Cl_2O_5 ?

1. chloroxide
2. dichlorine oxide
3. dichloride pentoxide
4. chlorine oxide
5. dichlorine pentoxide

011 (part 1 of 1) 10 points

The compound dinitrogen trioxide would have the chemical formula

1. N_3O_2 .
2. NO_3 .
3. N_2O_5 .
4. N_2O_3 .

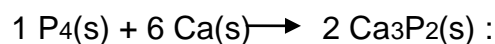
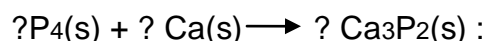
012 (part 1 of 1) 10 points

What type of reaction is
 $\text{AgNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$!
 $\text{NH}_4\text{NO}_3(\text{aq}) + \text{AgCl}(\text{s})$? note that a solid forms.

1. redox
2. precipitation
3. decomposition
4. single displacement

013 (part 1 of 1) 10 points

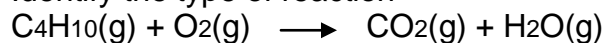
Balance the equation and identify the type of reaction for



1. 1; 6; 2 - redox
2. 2; 6; 2 -redox
3. 1; 6; 2 –synthesis (synthesis and not redox b/c compound is formed from elements)
4. 1; 6; 2 - decomposition
5. 2; 6; 2 - synthesis
6. 2; 6; 2 -decomposition

014 (part 1 of 1) 10 points

Identify the type of reaction



1. complete combustion
2. incomplete combustion
3. hydrocarbon synthesis
4. single displacement

015 (part 1 of 1) 10 points

Classify the reaction

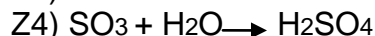
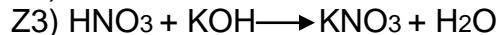
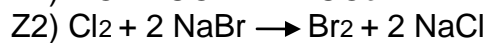


as one or more of the following:

- 1) combination; 2) decomposition; 3) redox;
 - 4) metathesis; 5) displacement
1. 2 only
 2. 2 and 3 only
 3. 1 and 3 only
 4. 5 and 3 only
 5. 2 and 4 only

016 (part 1 of 1) 10 points

Which response includes all of the following

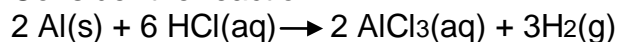


that are combination reactions, and none that are not combination reactions?

1. Z1, Z2 and Z4 only
2. Z1, Z3 and Z4 only
3. Z1 and Z4 only
4. Another reaction or another combination
5. Z1 and Z2 only

017 (part 1 of 1) 10 points

Consider the reaction

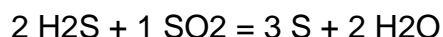


Which of the following is true?

1. Hydrogen is oxidized. (false, it's reduced)
2. HCl is the reducing agent. (false, it's oxidizing agent)
3. This is a displacement reaction.
4. This is a neutralization reaction.
5. Chlorine is reduced. (false, it's neither reduced nor oxidized here)

018 (part 1 of 1) 10 points

Vast quantities of elemental sulfur are produced from hydrogen sulfide by reacting it with sulfur dioxide (which itself can be prepared from hydrogen sulfide). When the Equation for the reaction between hydrogen sulfide and sulfur dioxide is properly balanced, the sum of the coefficients of the products and reactants is?



1. 10
2. 8
3. 4
4. 6

019 (part 1 of 1) 10 points

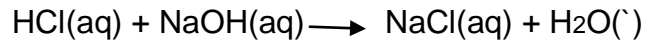
A ruptured tank in Anywhere, USA, spilled hydrochloric acid. Officials were reported to be concerned that the acid might mix with sodium hydroxide (caustic soda) in a neighboring tank and form a toxic gas. What toxic gas would be formed by mixing hydrochloric

acid with sodium hydroxide?

1. None (the chemicals would neutralize each other, however the heat from the reaction might make some of the HCl vaporize, and that would be a toxic gas!)
2. chlorine
3. chlorine dioxide
4. carbon dioxide

020 (part 1 of 1) 10 points

In the reaction



the precipitate is

1. NaCl.
2. There is NO precipitate.
3. HCl.
4. H₂O.
5. NaOH.

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