

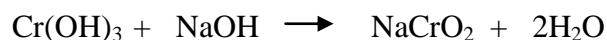
For answers, send email to: [admin@tutor-homework.com](mailto:admin@tutor-homework.com).

**Include file name:** Chemistry\_Worksheet\_0035

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1. 1.5 kg of fuel is burned on a natural gas grill during a barbecue. How much heat is released in kcal?
2. If you use a liter (1000mL) bottle of oil in your engine, how much heat is produced in kcal? In BTU? To solve this problem, you must know the density of motor oil is .82 g/mL.
3. If you eat a 6 oz. can (180 g) of tuna with ~.2 ppm Hg, how much mercury do you ingest?
4. A. What mass of NaCrO<sub>2</sub> can be obtained from the reaction of 7.40 g Cr(OH)<sub>3</sub> with 7.60g NaOH in the following reaction:



B. What is the limiting reactant?

5. Calculate moles of H<sub>2</sub> needed to form 8.5 moles HCl in the reaction  
 $\text{H}_2 + \text{Cl}_2 \longrightarrow \text{HCl}$ .
6. What is the pressure of a gas if 300ml at 0.8 atm and 37°C changes to 50°C and 555 ml?

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