

For each molecule listed below,

- (1) Determine the *total* number of *valence* electrons;
- (2) Write *Lewis electron dot structures*, including “resonance” forms, where applicable;
- (3) Predict the *molecular geometry*, including approximate *angles*, using VSEPR theory;
- (4) Determine the *polarity* of the molecule; if polar, show the *orientation* of the dipole.

NCCN (atoms bonded in order shown)	AsO ₄ ³⁻
Total number of valence electrons:	Total number of valence electrons:
Lewis electron dot structure(s):	Lewis electron dot structure(s):
Molecular geometry:	Molecular geometry:
Is the molecule polar? If YES, show orientation of dipole moment.	Is the molecule polar? If YES, show orientation of dipole moment.

For answers, send email to: admin@tutor-homework.com.

Include file name: Chemistry_Worksheet_0042

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