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**Include file name:** Chemistry\_Worksheet\_0087

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**What acid and base are neutralized to give sodium bromide salt?**

- a. HBr(aq) and NaOH(aq)
- b. HBr(aq) and NaBr(aq)
- c. HOH(aq) and NaBr(aq)
- d. HCl(aq) and NaOH(aq)
- e. none of the above

**If the concentration of hydronium in a solution is  $7.39 \times 10^{-5}\text{M}$ , what is the pH of the solution?**

**4.13**

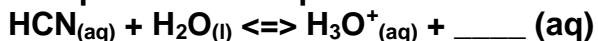
**In the following reaction, which reactant is a Bronsted-Lowry base?**  
 $\text{NaHCO}_3(\text{aq}) + \text{NaH}_2\text{PO}_4(\text{aq}) \rightarrow \text{Na}_2\text{HPO}_4(\text{aq}) + \text{H}_2\text{CO}_3(\text{aq})$

- a.  $\text{Na}_2\text{HPO}_4$
- b.  $\text{NaHCO}_3$
- c.  $\text{NaH}_2\text{PO}_4$
- d.  $\text{H}_2\text{CO}_3$
- e. none of the above

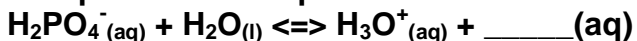
**If the concentration of hydroxide in a solution is  $3.57 \times 10^{-7}\text{M}$ , what is the pH of the solution?**

**7.55**

**Complete the  $K_a$  equation. Use the formatting used in OWL.**



**Complete the  $K_a$  equation. Use the formatting used in OWL.**



**If a light bulb in a conductivity apparatus glows dimly when testing a solution, which of the following is true?**

- a. The solution is slightly ionized.
- b. The solution is highly reactive.
- c. The solution is slightly reactive.
- d. The solution is highly ionized.
- e. none of the above

**What is the concentration of hydronium when the pH is 3.577**  
**0.000264 M**

**What is the term used interchangeably with hydrogen ion donor?**

- a. Arrhenius base
- b. proton donor
- c. Bronsted-Lowry base
- d. amphoteric
- e. none of the above

**Which of the following is a general property of an acidic solution?**

- a. feels slippery
- b. turns litmus blue
- c. tastes sour
- d. pH greater than 7
- e. none of the above

**What is the concentration of hydronium when the pOH is 9.36**  
 **$2.29 \times 10^{-5}$  M**