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SHOW UNITS W/EACH ANSWER; EX. 409 G

What concentration of NaOH solution was used to prepare 489.4 mL of 3.6 **M** NaOH if 0.039 L of the stock were used?

A solution is prepared by dissolving 51.41 g of KCl (74.551 g/mol) in enough water to make 367.3 mL of solution. What was the molarity of the solution?

28.35 g of 49.8 % Na_2CO_3 (105.989 g/mol) solution contains what mass of Na_2CO_3?

If a 50.0 mL sample of ammonium hydroxide is titrated with 33.5 mL of 4.8 **M** sulfuric acid to a methyl red endpoint, what is the molarity of the base?$2\ NH_4OH(aq) + H_2SO_4(aq) \rightarrow (NH_4)_2SO_4(aq) + 2\ H_2O(l)$

If 25.0 mL of2.55 **M** $Ca(OH)_2$ is titrated with 0.200 **M** HNO_3, what volume (mL) of nitric acid is required to neutralize the base?$2\ HNO_3(aq) + Ca(OH)_2(aq) \rightarrow 2\ Ca(NO_3)_2(aq) + 2\ H_2O(l)$

What is the term that refers to liquids that dissolve completely in one another?

- immiscible
- insoluble
- soluble
- miscible
- none of the above

What mass of NaCl (58.443 g/mol) is dissolved in 281 mL of a 4.4 **M** NaCl solution?

34.94 g of 53.8 % Na_2CO_3 (105.989 g/mol) solution contains what mass of solvent?

A solution of Na_2CO_3 (105.989 g/mol) is prepared by dissolving 14.7 g of the Na_2CO_3 in 94.1 g of water. What was the mass % of the solution?