

electron alternate additional points (Homework)

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1.

Write the full electron configuration ($1s^2 2s^2$, etc.) for each of the following elements. (Type your answer using the format $1s^2 2s^2 2p^6$ for $1s^2 2s^2 2p^6$.)

(a) krypton, $Z = 36$

(b) carbon, $Z = 6$

(c) neon, $Z = 10$

(d) chlorine, $Z = 17$

2.

Using the symbol of the previous noble gas to indicate the core electrons, write the valence shell electron configuration for each of the following elements. (Type your answer using the format $[\text{Ne}] 3s^2 3p^4$ for $[\text{Ne}] 3s^2 3p^4$.)

(a) vanadium, $Z = 23$

(b) scandium, $Z = 21$

(c) yttrium, $Z = 39$

(d) calcium, $Z = 20$

3.

How many $3d$ electrons are found in each of the following elements?

(a) vanadium, $Z = 23$

(b) iron, $Z = 26$

(c) scandium, $Z = 21$

(d) titanium, $Z = 22$