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Chemistry_Questions_0024

19.1 - Balance the following redox equations by the ion-electron method:

- a.) $\text{H}_2\text{O}_2 + \text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (in acidic solution)
- b.) $\text{Cu} + \text{HNO}_3 \rightarrow \text{Cu}^{2+} + \text{NO} + \text{H}_2\text{O}$ (in acidic solution)
- c.) $\text{CN}^- + \text{MnO}_4^- \rightarrow \text{CNO}^- + \text{MnO}_2$ (in base solution)
- d.) $\text{Br}_2 \rightarrow \text{BrO}_3^- + \text{Br}^-$ (in base solution)
- e.) $\text{S}_2\text{O}_3^{2-} + \text{I}_2 \rightarrow \text{I}^- + \text{S}_4\text{O}_6^{2-}$ (in acidic solution)

22.17 - Write the formulas for each of the following ions and compounds:

- a.) tetrahydodoxozincate(II)
- b.) pentaaquochlorochromium(III)
- c.) tetrabromocuprate(II)
- d.) ethylenediaminetetraacetatoferate(II).

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