

nomenclature practice test (Homework)

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Chemistry_Questions_0129

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1.

Name each of the following binary compounds of nonmetallic elements.

- (a) IF_5
- (b) AsCl_3
- (c) SeO
- (d) XeF_4
- (e) NI_3
- (f) B_2O_3

2.

Name each of the following binary compounds, using the periodic table to determine whether the compound is likely to be ionic (containing a metal and a nonmetal) or nonionic (containing only nonmetals).

- (a) CaH_2
- (b) OCl_2
- (c) XeF_4
- (d) CIF
- (e) Ca_3N_2
- (f) Ag_2S

3.

Write the formula for each of the following phosphorus-containing ions, including the overall charge of the ion. (Type your answer using the format H_3PO_4 for H_3PO_4 , $(\text{NH}_4)_2\text{CO}_3$ for $(\text{NH}_4)_2\text{CO}_3$, $[\text{NH}_4]^+$ for NH_4^+ , and $[\text{Ni}(\text{CN})_4]^{2-}$ for $\text{Ni}(\text{CN})_4^{2-}$.)

- (a) dihydrogen phosphate
- (b) dihydrogen phosphide
- (c) dihydrogen phosphite
- (d) phosphate

4.

Name each of the following acids.

- (a) HI
- (b) HBrO_4

- (c) HNO_2
- (d) HCl
- (e) HOI
- (f) H_2Se
- (g) HBr
- (h) H_2SO_3
- (i) HClO_3

5.

Write the formula for each of the following substances. (Type your answer using the format CH_4 for CH_4 and $(\text{NH}_4)_2\text{CO}_3$ for $(\text{NH}_4)_2\text{CO}_3$.)

- (a) rubidium hydroxide
- (b) tin(IV) acetate
- (c) nitrous acid
- (d) ammonium dichromate
- (e) lithium sulfate
- (f) calcium chlorate
- (g) hypochlorous acid
- (h) hydrotelluric acid

- (i) nitric acid
- (j) potassium sulfate
- (k) cesium perchlorate
- (l) lithium peroxide

6.

Write the formula for each of the following substances. (Type your answer using the format $\text{Cr}(\text{CN})_3$ for $\text{Cr}(\text{CN})_3$.)

- (a) potassium sulfate
- (b) cesium perchlorate
- (c) barium acetate
- (d) calcium bromide
- (e) iron(II) oxide
- (f) rubidium hydroxide
- (g) calcium chlorate
- (h) strontium nitrate
- (i) ammonium dichromate
- (j) dinitrogen tetroxide

(k) cobalt(III) nitrate

(l) lithium peroxide

7.

Name the following compounds. (Type your answer using the format iron(II) for Fe^{2+} .)

(a) Cr_2S_3

(b) $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$

(c) B_2O_3

(d) $\text{Fe}(\text{NO}_3)_3$

(e) CoCl_3

(f) $\text{Cu}(\text{MnO}_4)_2$

(g) LiHCO_3

8.

Write the formula of each of the following ionic substances. (Type your answer using the format CO_2 for CO_2 .)

(a) lithium perchlorate

(b) barium peroxide

(c) cesium sulfite

(d) cuprous iodide

(e) chromic fluoride

(f) sodium dihydrogen phosphate